

Supplementary Information

Printable homocomposite hydrogels with synergistically reinforced molecular-colloidal networks

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Supplementary Video 1 | Synergistic effect in mixed SA-SDC compositions for 3D printing. A solution of 1.5 wt.% molecular SA in H₂O and a suspension of 1.5 wt.% SDC are both in fluid-like state. Once they are mixed in 1:1 ratio to yield a 1.5 wt.% homocomposite mixture, they form a gel-like extrudable system even prior to cross-linking of the molecular alginate network.

Supplementary Video 2 | 3D printing via extrusion of a layered design with 1.5 wt.% SA homocomposite gel system (0.75% SDC, 0.75% CMH).

Supplementary Results and Discussion

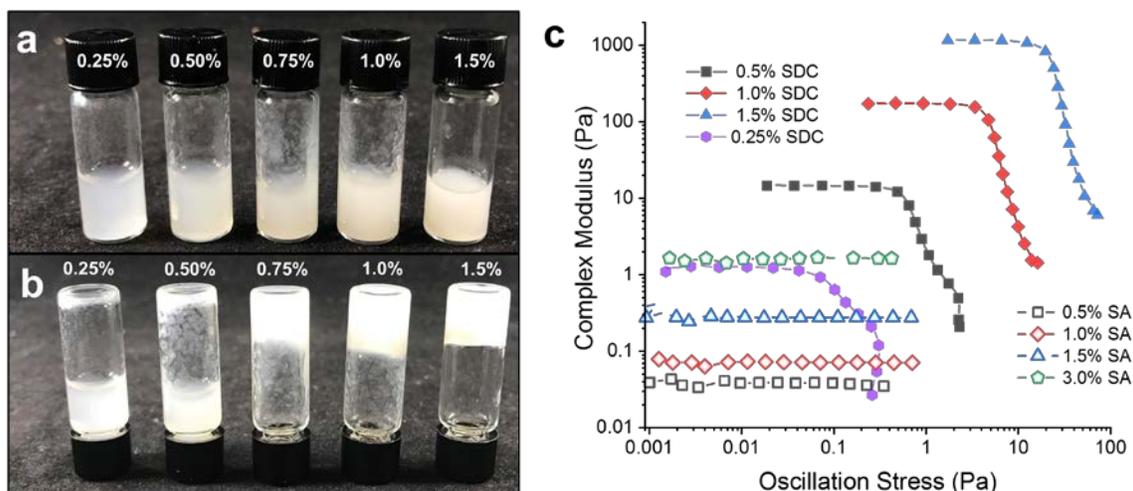
Alginate Hydrogel Soft Dendritic Colloid Fabrication

In order to make the nanofibrous reinforcement for HHGs presented herein, we synthesized sodium alginate (SA) soft dendritic colloids (SDCs) with hierarchical, branched micro- and nanofibers. SA SDCs are fabricated using shear-driven polymer precipitation within a nonsolvent flow of an aqueous CaCl₂ solution. Upon injection of the SA solution into the turbulently sheared precipitation medium, the solution is randomly and vigorously stretched and templated by the hierarchical turbulent eddies of the Ca²⁺ rich medium. Concurrent with the multidirectional

deformation of the polymer solution by the turbulent flow, the Ca^{2+} ions rapidly diffuse into the stretched SA solution droplet. The simultaneous droplet elongation by turbulent eddies and Ca^{2+} diffusion result in the precipitation of the alginate into highly branched and fibrous SA SDCs. While the morphology remains the same, these SDCs are fundamentally different from those of our previous report. Herein, we rely upon the rapid, ionic crosslink reaction between alginate and Ca^{2+} as the precipitation mechanism to form water-swollen and cross-linked SA SDC fibers. Contrarily, our previous report utilized nonsolvent-induced precipitation, which resulted in similarly structured, but more densely packed polymer fibers. While the precipitation mechanisms of the two materials may differ, the resulting particulates of both processes possess the characteristic SDC morphology. We hypothesize that parameters such as temperature, pH, and exposure to light that can induce rapid precipitation in other polymer systems can also be utilized as precipitation mechanisms under shear to obtain SDCs from a vast library of materials.

Characterization of Alginate SDCs

SA SDC suspensions were washed with deionized water to remove redundant Ca^{2+} ions using centrifugation following by decanting and resuspension of the SDC pellet in deionized water. After rinsing, a dilute suspension of SDCs (< 0.001 wt.%) in ethanol was prepared and spin-coated on a silicon wafer and a glass slide to observe individual SA SDCs with a field-emission scanning electron microscopy and optical microscopy (Fig. 1b). SDCs can often be particulates spanning hundreds of microns in total diameter, however, the fibers comprising the greater particulate branch into smaller, nanofibers whose networking is dominated by physical interactions and fiber entanglements.

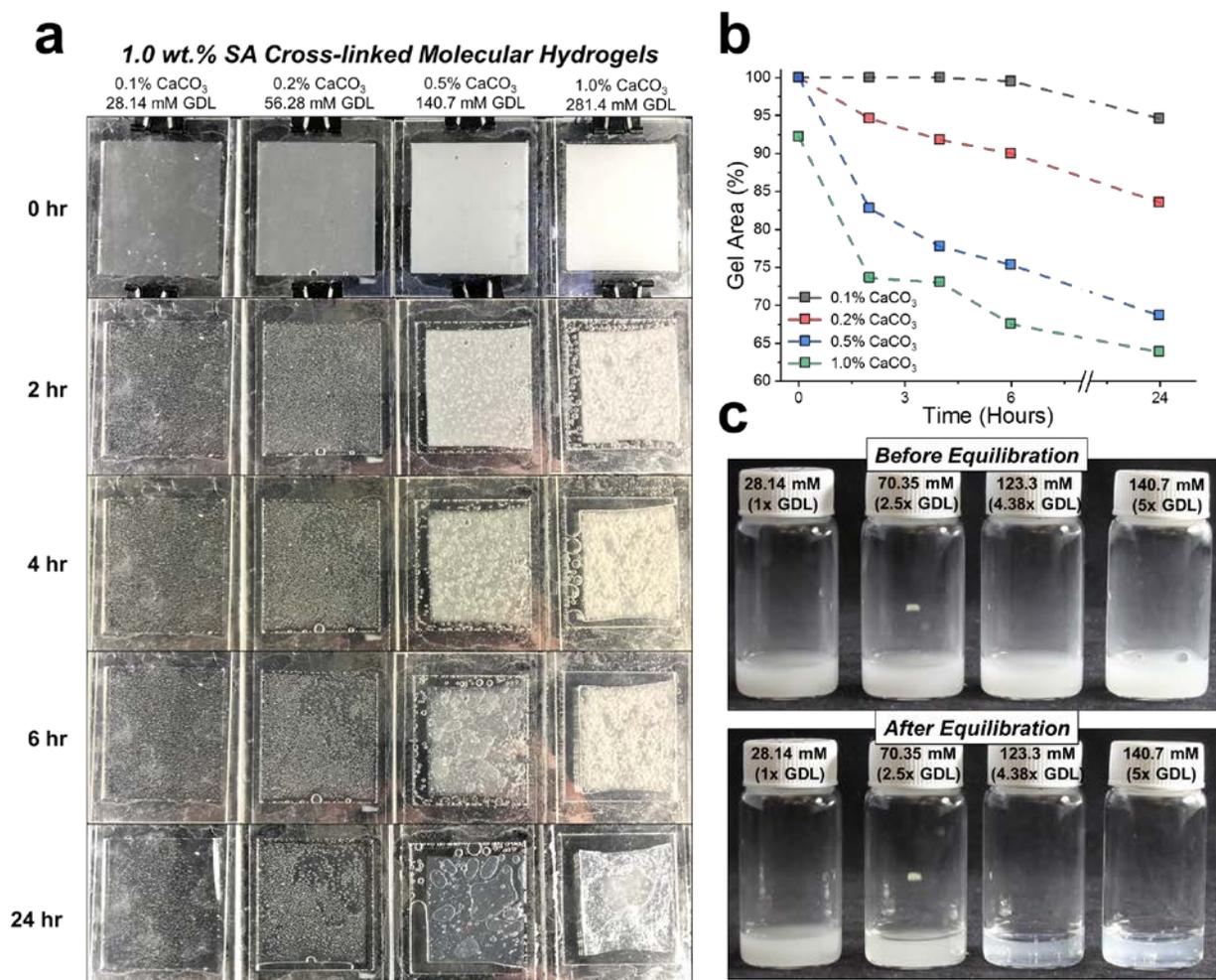


Supplementary Figure 1 | Gelation of SA SDC suspensions. Pictures of aqueous SA SDC suspensions at concentrations ranging 0.25 – 1.5 wt.% **a** before and **b** after inversion of the vials indicating apparent yield stress and gel-like behavior with increasing concentration. **c** Amplitude sweeps performed on dissolved SA solutions (unfilled) and aqueous SA SDC suspensions (filled) of varied concentrations showing yield stress behavior in dilute SDC suspensions.

The pictures of the SDC solutions in Supplementary Fig. 1 show that while suspensions of SA SDCs show gel-like behavior when analyzed using a frequency sweep down to concentrations of 0.25 wt.%, these dilute suspensions flow upon inversion of their vials until a concentration between 0.50 – 0.75 wt.%. The impressive network-building capabilities by SDC suspensions are evidenced by their yield stress behavior at dilute concentrations as low as 0.25 wt.% (Supplementary Fig. 1c), while non-cross-linked SA solutions show Newtonian behavior up to concentrations of 3%. Interestingly, the SDC suspension at just 0.25% has a complex modulus nearly matching that of the 3% SA solution containing 12× the alginate.

To determine the CaCO_3 loading capacity of the SA CMHs, a series of 1.0 wt.% SA gels were prepared with increasing CaCO_3 content and the GDL added to the gels was raised proportionally with the CaCO_3 content (Supplementary Fig. 2a). These gels showed significant syneresis with increased CaCO_3 and GDL content, which was quantified by using image analysis

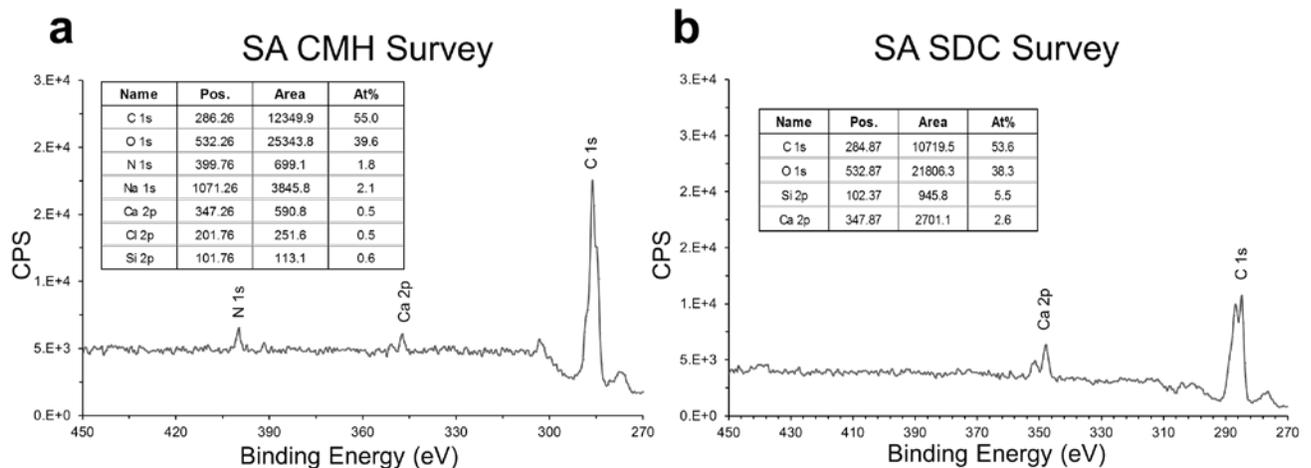
to measure the decrease in the area of the gels over time within their 50 mm × 44 mm × 1.5 mm acrylic molds (Supplementary Fig. 2b). The SA CMHs containing high CaCO₃ loading (0.5 wt.% and 1.0 wt.%) do not become completely transparent after 24 hours indicating incomplete CaCO₃ dissolution and lose significant volume by syneresis. Next, a set of gels containing 1.0 wt.% SA, 0.5% CaCO₃, and varying concentrations of added GDL were prepared. It can be seen visually that even up to 140.7 mM GDL, the mixture retains its opacity, indicating incomplete dissolution of the CaCO₃ nanoparticles (Supplementary Fig 2c).



Supplementary Figure 2 | Syneresis of SA CMHs with increased CaCO₃ content. **a** Pictures of gel syneresis prepared in 50 mm × 44 mm × 1.5 mm sealed acrylic molds and **b** the area of the gels over time measured by image analysis in ImageJ of the area loss of the gels over time. **c** Pictures of 1.0% SA, 0.5% CaCO₃ mixtures immediately after addition of 28.14 – 140.7 mM GDL.

Elucidating the Role of Calcium in SDC Gelation

To determine the level of calcium incorporated into the fibers of the SA SDCs during fabrication, a 1.0 wt.% suspension of SA SDCs was repeatedly washed with deionized water to remove excess ions from solution and 8 mL of the suspension was deposited in an aluminum dish and dried at 50°C for 24 hours. A 1.0 wt.% molecular alginate gel was also prepared with 0.11 wt.% CaCO₃ NPs and 28.14 mM GDL. This mixture was then deposited in an aluminum weighing dish, allowed to gel, and placed in a 50°C oven for 24 hours to dry to produce a thin, cross-linked alginate film of 10 wt.% CaCO₃. These films were placed under vacuum for 2 days and analyzed by X-ray photoelectron spectroscopy (XPS, Supplementary Fig. 3). The atomic percentage of calcium in the CMH film was determined to be 0.5%, while that of the SDC film was found to be 2.6%, indicating that the SDCs contain approximately 5× the amount of calcium than the molecular alginate film at the prepared concentration.



Supplementary Figure 3 | X-ray photoelectron spectroscopy analysis of dried molecular alginate and SDC films. a XPS spectrum of a dried molecular alginate film prepared with by drying a 1 wt.% SA, 0.10 wt.% CaCO₃ NPs, 28.14 mM GDL CMH into a film. **b** XPS spectrum of a dried SA SDC film after repeated washing an aqueous SDC suspension with deionized water and drying.

With SA SDCs containing appr. 5× the calcium content of the CMH, it is then possible to calculate the concentrations in [mM] of each constituent of the gel (Supplementary Table 1).

With the maximum HHG tensile stiffness resulting from HHGs with a composition of 0.125% SDC/0.875% (Fig. 3), it can then be estimated the calcium content of this HHG is appr. 15.2 mM, or 1.5× that of the pure CMH.

Supplementary Table 1. Estimated equilibrium compositions of 1.0% SA CMH/SDC HHGs in [mM].

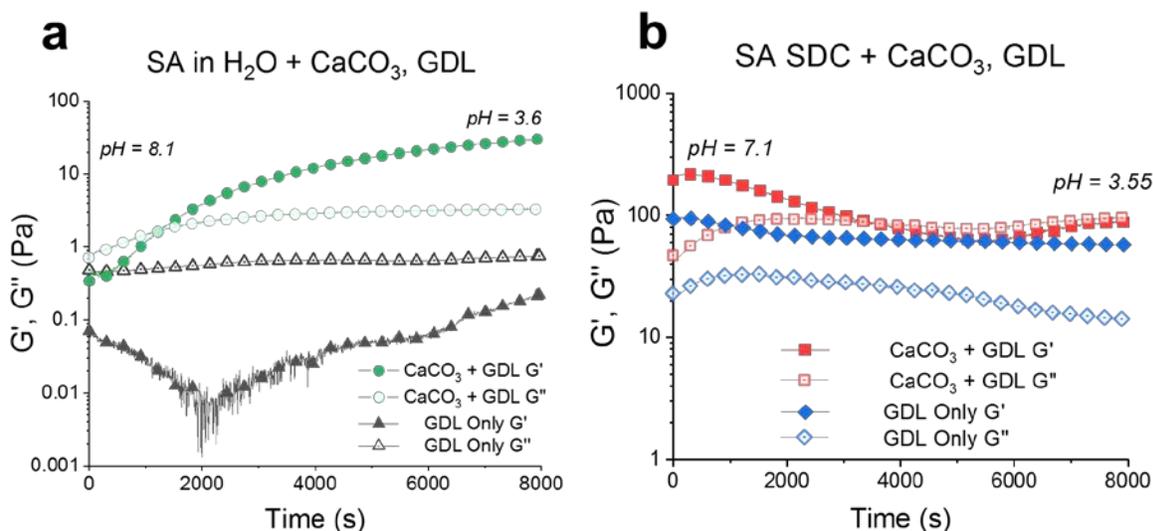
SA CMH/SDC Homocomposite Compositions (mM)

wt.% SDC	0.000	0.125	0.250	0.500	0.750	1.000
wt.% CMH	1.000	0.875	0.750	0.500	0.250	0.000
Total [Ca ²⁺] (mM)	9.991	15.236	20.482	30.972	41.463	51.953
[GDL] (mM)	28.143	24.625	21.107	14.071	7.036	0.000

Rheological Time Sweep Analysis of SA CMHs and HHGs

To ensure that the ionization of CaCO₃ NPs by the GDL-mediated pH drop was solely responsible for the gelation seen in CMH and homocomposite hydrogel samples, a series of time sweep experiments were performed in which only GDL as well as CaCO₃ and GDL were added to pure SA CMH and SDC samples. When the combination of CaCO₃ NPs and GDL were added to a SA solution, the sample saw gelation as expected. Without the presence of the CaCO₃ NPs, the SA solution did not see any gelation despite the drop in the solution pH from 8.1 to 3.6 after completing equilibration (Supplementary Fig. 4a). The same experiments were performed on a 0.67 wt.% SA SDC suspension in deionized water and it was found that the addition of CaCO₃ did increase the storage and loss moduli of the suspension slightly initially, however, the dissolution of those CaCO₃ NPs over time by GDL hydrolysis led to the weakening of these suspensions and eventually, fluid-like behavior with the loss modulus eclipsing the storage modulus

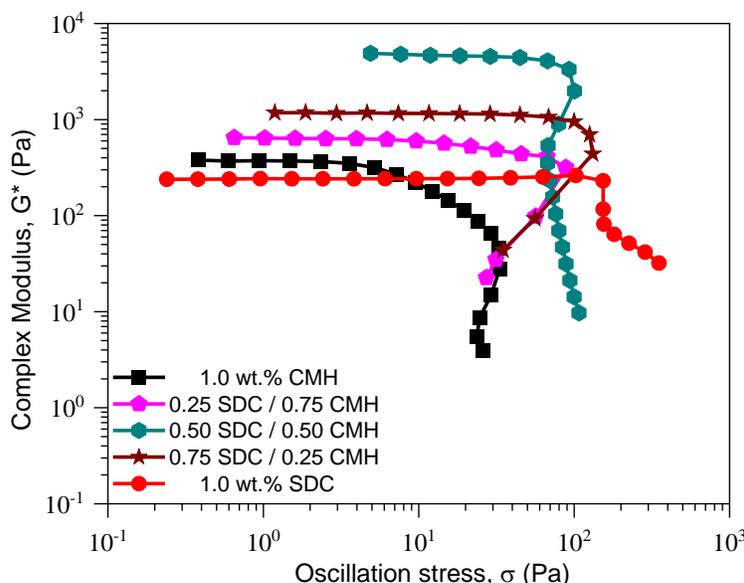
(Supplementary Fig. 4b). This result indicates that the SA SDCs are unable to be further cross-linked in suspension by the introduction of Ca^{2+} ions *in situ*. Furthermore, the addition of solely GDL to the SDC suspension resulted in the weakening of the suspension over time. The combination of these results with the CMH and SDC systems shows that the greatly increased moduli seen in the homocomposite hydrogels is not a byproduct of any GDL-mediated pH drop or tertiary cross-linking phenomena between the SDCs and the CaCO_3 -GDL combination, but instead, is a result of the interplay between the molecular alginate network of the CMH and the colloidal alginate network of the SDCs.



Supplementary Figure 4 | Time sweeps of SA SDCs and molecular alginate solutions with addition of CaCO_3 and GDL. **a** Time sweeps of 1.0 wt.% SA SDC suspensions with and without 0.1% CaCO_3 NPs following addition of 28.14 mM GDL. **b** Time sweeps of 0.67 wt.% SA SDC suspensions with and without 0.067% CaCO_3 NPs following addition of 18.77 mM GDL. Both graphs show the pH of the suspensions and solutions before and after GDL addition and equilibration.

Rheological analysis of the 1.0 total wt.% HHGs following equilibration (Supplementary Fig. 5) shows the maximum complex modulus of the 1.0 total wt.% equilibrated alginate HHGs

resulted at a 1/1 SDC/CMH ratio, despite the maximum stiffness being obtained at lower SDC concentrations. A maximum complex modulus being obtained at 1/1 a SDC/CMH formulation is seen again in Figure 5a with the 1.5 total wt.% HHG mixtures prior to GDL addition.

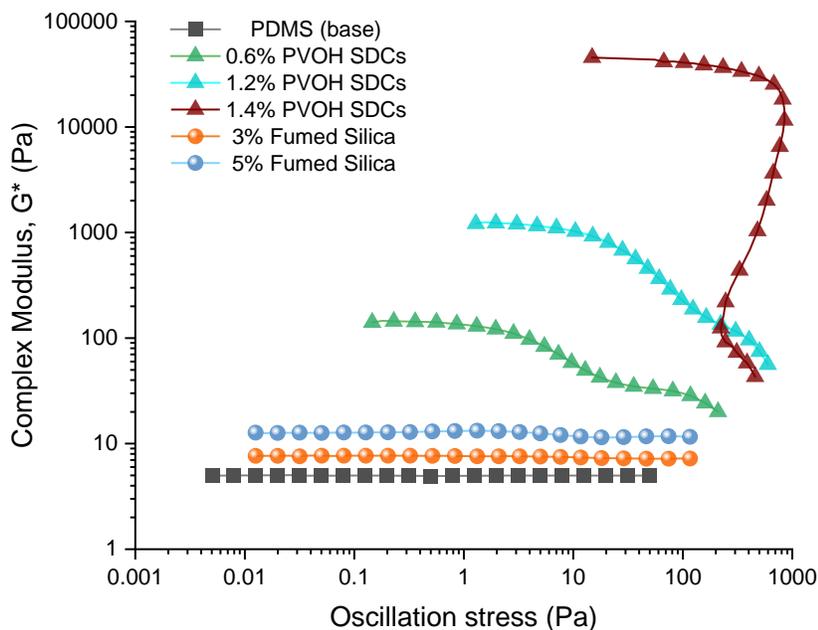


Supplementary Figure 5 | Amplitude sweeps of 1.0 wt.% total alginate HHG formulations.

Cross-linked 1% HHGs of varying composition performed 2 hours after GDL addition at a frequency of 6.23 rad/s. These results show a maximum complex modulus for equilibrated HHGs is obtained at a 1/1 ratio of SDC/CMH.

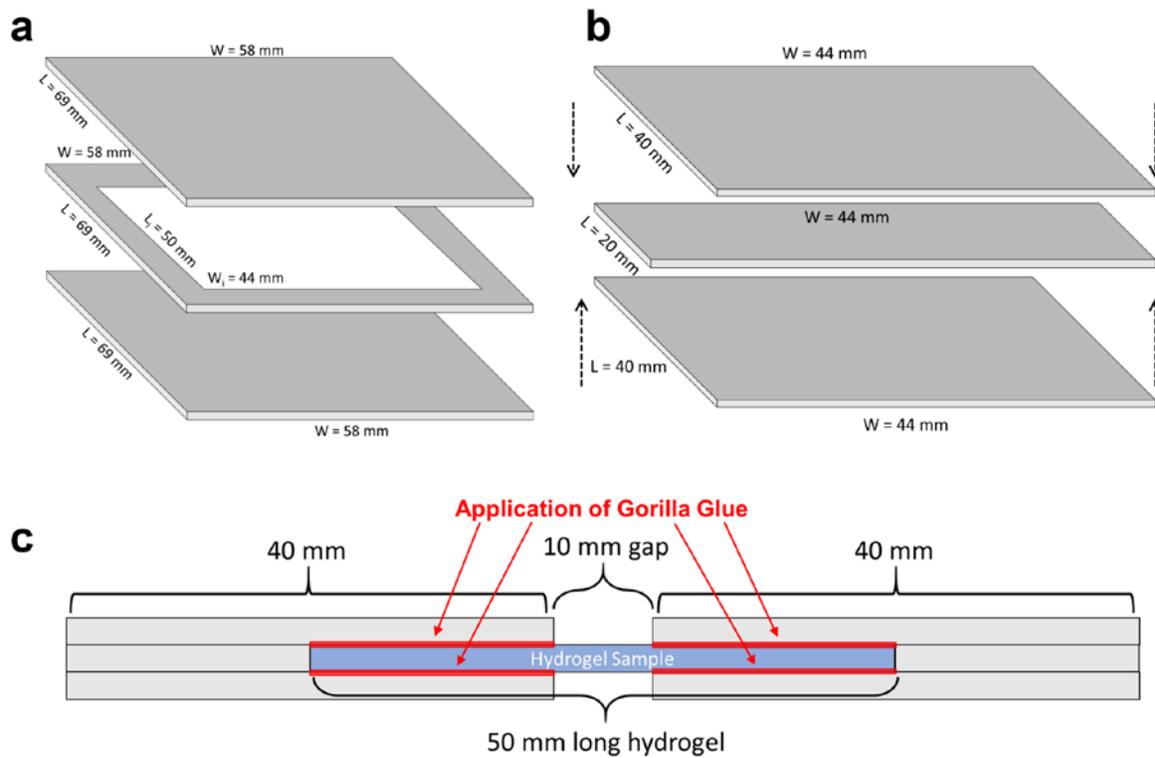
In Supplementary Fig. 6, an example of SDCs utilized as a yield stress additive in Sylgard 184 PDMS base is shown. The amplitude sweeps of SDC/PDMS composites show a 4-decade increase in the complex modulus with just 1.4 wt.% polyvinyl alcohol (PVOH) SDCs compared to the pure PDMS base. The amplitude sweeps of fumed silica (Aerosil R202) in show that the SDCs at just 0.6 – 1.4 wt.% far outperform the fumed silica in terms of viscoelastic modification of the continuous, liquid PDMS phase even when the silica content is increased to four-fold that of the SDCs. This result highlights the unrivaled efficiency with which SDCs can form physical

networks within a secondary liquid medium, as fumed silica is an extremely common thickening agent for inducing thixotropy and yield stresses through a percolated, colloidal network.



Supplementary Figure 6 | SDCs as universal composite yield stress additives. Amplitude sweeps of polyvinyl alcohol (PVOH) soft dendritic colloids compared with fumed silica dispersed in PDMS base (Sylgard 184) showing yield stress emergence at low SDC concentrations (w/w).

To prepare the HHGs with desired shape for tensile analysis and the syneresis measurements, laser cut acrylic sheets were used to make curing molds of desired dimensions for hydrogel curing and grippers for tensile analysis. A thin layer of cyanoacrylate adhesive (Gorilla Glue Gel) was applied at the gel-gripper interface and allowed to cure for 30 minutes prior to tensile analysis (Supplementary Fig. 7).



Supplementary Figure 7 | Schematic of the laser-cut acrylic **a** mold used to fabricate hydrogel sample for tensile analysis, **b** grippers used to hold the hydrogel during tensile analysis, and **c** the assembled grippers chemically bonded to the hydrogel sample with clamping zones indicated. All acrylic sheets were 1.59 mm thick and solvent-welded together using dichloromethane.